

DEPARTMENT OF COMMERCE AND LABOR
BUREAU OF STANDARDS
S. W. STRATTON, Director

THE DETERMINATION OF TOTAL SUL-
PHUR IN INDIA RUBBER

BY

C. E. WATERS, Associate Chemist

and

J. B. TUTTLE, Assistant Chemist

Bureau of Standards

[JULY 19, 1911]

REPRINT NO. 174

(FROM BULLETIN OF THE BUREAU OF STANDARDS, VOL. 8, NO. 3)



WASHINGTON
GOVERNMENT PRINTING OFFICE

1912

12-35432

DEPARTMENT OF COMMERCE AND LABOR

BUREAU OF STANDARDS

S. W. STRATTON, Director

THE DETERMINATION OF TOTAL SUL-
PHUR IN INDIA RUBBER

BY

C. E. WATERS, Associate Chemist

and

J. B. TUTTLE, Assistant Chemist

Bureau of Standards

[JULY 19, 1911]

REPRINT NO. 174

(FROM BULLETIN OF THE BUREAU OF STANDARDS, VOL. 8, NO. 3)



WASHINGTON
GOVERNMENT PRINTING OFFICE

1912

TS1892
W3
C16

U. S. DEPT. OF JUSTICE
FEB 22 1912

12-35452

THE DETERMINATION OF TOTAL SULPHUR IN INDIA RUBBER

By C. E. Waters and J. B. Tuttle

What may be still described as the usual method for the determination of total sulphur in india rubber is the one first published by Henriques.¹ The details of this method are too well known to require description here. In more recent years other methods have been advocated. Alexander² used sodium peroxide to decompose the nitrogen peroxide addition-product of rubber. In the same year Esch³ recommended the use of Eschka's mixture and procedure for the determination of sulphur in coal. He also stated that the sodium-peroxide method gives good results.

Wagner⁴ published a slight modification of the method of Henriques, stating that much sulphur is lost by volatilization. He, therefore, made the nitric-acid solution alkaline with sodium hydroxide, transferred to a nickel crucible, added sodium carbonate and then evaporated to dryness. The oxidation was carried to completion by heating in an air-bath.

Pontio⁵ fused with manganese peroxide and a mixture of sodium and potassium carbonates. The results were about 0.1 per cent lower than by the method of Henriques. For the free sulphur⁶ he extracted with absolute alcohol, distilled off the solvent, oxidized with alkaline hydrogen peroxide, evaporated to dryness and fused in a silver crucible.

¹ *Z. angew. Chemie*, **12**, p. 902; 1899.

² *Gummi-Ztg.*, **18**, p. 729; *Z. angew. Chemie*, **17**, p. 1799; 1904.

³ *Chem.-Ztg.*, **28**, p. 200; 1904.

⁴ *Gummi-Ztg.*, **21**, p. 552; *Chem. Abstr.*, **1**, p. 1327; 1907.

⁵ *Caoutchouc et Gutta-Percha*, **6**, p. 2751; *Chem. Techn. Rep.*, **1909**; 372.

⁶ *Caoutchouc et Gutta-Percha*, **5**, p. 2194; *Chem. Abstr.*, **2**, p. 3412; 1908.

A distinct departure from the usual methods is due to Hinrichsen,⁷ who oxidizes electrolytically in the presence of concentrated or fuming nitric acid.

Finally, Hübener⁸ devised a method intended to exclude insoluble mineral sulphates. The sample is boiled in a flask with concentrated nitric acid for some time, most of the acid is evaporated off on the steam-bath, and the oxidation completed by means of bromine and water.

One of the present writers, having frequent occasion to determine total sulphur in rubber, over a year ago made a number of comparative tests of different variations of the method of Henriques. The results obtained with two samples of rubber are given below (I-V). In all cases, 0.50 g of rubber was taken. All fusions were made over a flame of gasoline-air gas. The results are given as percentages of sulphur. All reagents were tested, and no determinations have been omitted.

I. Warmed $2\frac{1}{2}$ hours in covered crucible with 25 cc conc. HNO_3 , allowed to stand 36 hours, evaporated nearly to dryness, added Na_2CO_3 , KNO_3 mixture and fused as usual.

Sample	1	1	2	2
Sulphur	3.39	3.44	3.26	3.22

II. Added HNO_3 and 1 cc Br, let stand 36 hours without preliminary heating, evaporated, etc., as usual.

Sample	1	1	2	2
Sulphur	3.40	3.47	3.39	3.27

III. The same as II, but allowed to stand only 1 hour, heated with cover for 2 hours, evaporated and fused as usual.

Sample	1	1	2	2
Sulphur	3.31	3.35	3.04	3.09

IV. Only HNO_3 added, digested at once on the steam-bath for 2 hours, evaporated and fused.

Sample	1	1	2	2
Sulphur	3.17	3.43	3.06	2.93

⁷ Chem.-Ztg., 33, p. 735; 1909.

⁸ Gummi-Ztg., 24, pp. 213-214; Analyst, 35, pp. 266-267; 1910.

V. Treated with 1 cc Br and 5 cc H₂O, let stand over night without heating, next morning evaporated off the H₂O, added HNO₃, digested, evaporated and fused.

Sample	1	1	2	2
Sulphur	3.71	3.65	3.37	3.38

In an attempt to obtain satisfactory results without fusion, some determinations were made some months later, without a knowledge of Hübener's paper. Half-gram portions of a sample of medium hard rubber were digested with nitric acid in flasks covered with watch glasses. In some cases bromine was added after the digestion with acid and, after standing half an hour, water was added and the flasks heated on the steam-bath. Finally the volume was brought to about 175 cc, the solution heated, filtered, and a little sodium hydroxide added to the filtrate and wash-water. This was then evaporated to dryness, adding a little hydrochloric acid toward the end, taken up with very dilute hydrochloric acid, filtered and barium sulphate precipitated as usual. The results follow:

VI. Treated with HNO₃ alone.

Sulphur	7.76	7.51	7.68	7.96
---------	------	------	------	------

VII. Treated with HNO₃, followed by Br.

Sulphur	7.62	7.51	7.93	7.87	7.76
---------	------	------	------	------	------

All the precipitates obtained under VI and VII contained much lead.

After the method of Hübener was called to our attention some determinations were made on a sample of hard rubber containing no barium.

VIII. Hübener's method.

Sulphur	4.79	3.91	5.23	4.02	4.31	4.13
---------	------	------	------	------	------	------

It is evident that widely different amounts of sulphur must have been retained in the insoluble residue in the form of lead sulphate.

IX. Total sulphur by method of Henriques.

Sulphur	8.65	8.70
---------	------	------

X. Treated with HNO₃, followed by Br and H₂O and fused as usual.

Sulphur	8.63	8.62	8.77	8.80	8.72	8.80
---------	------	------	------	------	------	------

It has recently been claimed by van't Kruys⁹ that when an excess of calcium chloride over the amount of sulphuric acid is present, only calcium sulphate is carried down with the barium sulphate, and the calcium salt can be converted into barium sulphate by digestion with strong hydrochloric acid, or aqua regia, and barium chloride. Several determinations were made to test this suggestion.

XI. Preliminary treatment as under X, subsequent treatment as suggested by van't Kruys.

Sulphur 8.73 8.76 8.52 8.46 8.74 8.78 8.75 8.77

At this point, joint analyses of a fairly large sample of rubber were carried out by the present writers.

XII. Hübener's method. The sulphur in the insoluble residue was determined by fusion with soda-salt-peter mixture, extracting the melt with water, etc., as usual:

S in original filtrate 0.91 0.91 0.68 0.79 0.82 0.99 1.52 1.15 1.24

S in insoluble residue 2.35 2.43 2.54 2.69 2.53 2.36 1.95 2.31 2.21

Total sulphur found 3.26 3.34 3.22 3.48 3.35 3.35 3.47 3.46 3.45

All the precipitates of barium sulphate from the original filtrates were found to contain lead when tested with dilute ammonium sulphide.

XIII. Treated with HNO_3 , allowed to stand over night, the acid driven off on the steam-bath, 1 cc Br and 10 cc H_2O added; then the H_2O and excess of Br driven off by heating. The residue was mixed with soda and salt-peter and fused as usual.

Sulphur 3.41 3.21

XIV. The same as XIII, but the HNO_3 not driven off before adding Br.

Sulphur 3.60 3.63 3.58 3.63 3.57

XV. The same as XIII, but treatment with Br omitted.

Sulphur 3.29 3.49 3.35 3.38 3.55 3.43 3.36 3.58 3.56

XVI. Treated first with bromine and water, allowed to stand over night without heating, then Br and H_2O driven off on steam-bath, treated with HNO_3 , etc., and fused.

Sulphur 3.45 3.47 3.48 3.53 3.49

XVII. The same as XVI, but excess Br and H_2O not driven off before adding HNO_3 .

Sulphur 3.59 3.47 3.64

⁹ Zs. anal. Chem., 49, p. 393; 1910.

XVIII. The method of Henriques, except that the HNO_3 was saturated with Br.

Sulphur 3.66 3.65 3.62 3.73 3.68 3.65 3.69 3.63 3.66 3.71 3.62

XIX. The same as XVIII, but followed by the treatment suggested by van't Kruys.

Sulphur 3.69 3.73 3.76 3.75

In order to obtain a definite idea of the variations caused by differences in the preliminary treatment and in the conditions under which the barium sulphate is precipitated, a very dilute solution of sulphuric acid was made. In each of the following determinations, a 25-cc portion was taken. The weights of the barium sulphate found were calculated as percentages of sulphur in 0.50 g of rubber, in order that the results might be more readily compared with the determinations above.

XX. Direct precipitation with BaCl_2 . The last two determinations were made with the addition of 2 cc of 1:1 HCl, the first six without adding HCl.

Sulphur 3.11 3.11 3.11 3.11 3.11 3.11 3.10 3.11

XXI. Evaporated off the water from 25 cc of the dilute H_2SO_4 , added the soda-salt-peter mixture and fused as usual.

Sulphur 3.12 3.15 3.12

XXII. Like XXI, but added CaCl_2 to the solution of the melt before precipitating BaSO_4 , and treated the latter according to van't Kruys.

Sulphur 3.22 3.18 3.19

XXIII. Like XXII, but did not digest the precipitated BaSO_4 , nor evaporate the filtrate to recover traces of dissolved BaSO_4 .

Sulphur 3.13 3.12 3.11

XXIV. Like XXI, but did not fuse. The solution was acidified with HCl.

Sulphur 3.17 3.19 3.16

XXV. Added 250 cc H_2O and 10 cc conc. HCl to 25 cc dilute H_2SO_4 , added BaCl_2 , digested two hours, poured off the supernatant liquid, digested the BaSO_4 with 1 cc of 10 per cent BaCl_2 and 15 cc HCl (1:1). Diluted, filtered, evaporated the combined filtrates to dryness in platinum, took up with 50 cc of slightly acidified H_2O and collected the slight residue on the same filter.

Sulphur 3.12 3.10 3.11

XXVI. Exactly neutralized NaOH solution with the dilute H_2SO_4 , using phenolphthalein as indicator. After each addition of acid, the solution was heated in a platinum dish until the pink color no longer reappeared. Then evaporated to dryness, ignited gently, and weighed the Na_2SO_4 .

H_2SO_4 used (cc)	75.56	75.55
Na_2SO_4 found (g)	.2082	.2080
BaSO_4 equivalent to Na_2SO_4 (g)	.34197	.34164
BaSO_4 equivalent to 25 cc H_2SO_4 (g)	.1131	.1130
Sulphur (calc. on 0.5 g rubber)	3.11	3.11

XXVII. The same as XXVI, but used NaHCO_3 instead of NaOH.

H_2SO_4 used (cc)	45.01	43.48
Na_2SO_4 found (g)	.1240	.1194
BaSO_4 equivalent to Na_2SO_4 (g)	.20367	.19612
BaSO_4 equivalent to 25 cc H_2SO_4 (g)	.1131	.1128
Sulphur (calc. on 0.5 g rubber)	3.11	3.10

In order to test the completeness of the oxidation of sulphur by means of the nitric acid-bromine mixture, the following determinations were carried out.

XXVIII. Powdered sulphur crystals, digested in the cold with 20 cc HNO_3 and an excess of Br. Finally added 20 cc H_2O and heated on the steam-bath for about two hours. Then evaporated nearly to dryness, took up with water, and precipitated with BaCl_2 .

Sulphur taken (g)	0.0483	0.0395	0.0561
Sulphur found (g)	.0481	.0399	.0563

XXIX. Powdered sulphur crystals treated at the same time as some of the samples of rubber. The exact methods are referred to in the table, the Roman numerals indicating the method employed.

Method	I	II	III	IV	V
Sulphur taken (g)	0.0528	0.0619	0.0494	0.0646	0.0411
Sulphur found (g)	.0485	.0590	.0479	.0595	.0414
Sulphur, per cent	91.84	95.33	96.91	92.13	100.67

In the determinations by methods I to IV, part of the sulphur was not attacked by the nitric acid, nor by the sodium carbonate added before making the fusion. Part, at least, of this unattacked sulphur was seen to burn when the fusion was made.

As stated above (XII), the barium sulphate precipitates representing soluble sulphates, etc., in the Hübener method, were found to contain lead. Lead sulphate dissolves slightly, and is, besides, partially decomposed by water, hydrobromic and nitric acids, etc.¹⁰ In order to get an idea of the amount of barium sulphate to be expected to result from the decomposition and solution of lead sulphate under the conditions of the Hübener method, some determinations were made. Lead sulphate was first prepared by precipitation from a hot, dilute nitric-acid solution of lead nitrate by means of a hot, dilute solution of sulphuric acid. It settled rapidly as a coarse-grained powder, which was washed by decantation with hot water, then in a Gooch crucible with hot water, followed by strong alcohol. It was then dried in an air-bath.

In the first experiments it was treated with hot water, and the amount of barium sulphate precipitated from the filtrate was calculated as percentage of sulphur in 0.50 g rubber.

XXX. Washed 0.200-gram portions of PbSO_4 on filters. Each time 250 cc hot water was used. The filtrates were slightly acidified with HCl , and precipitated with BaCl_2 .

Sulphur 0.24 0.25

These precipitates contained only traces of lead.

XXXI. Treated 0.200-gram portions of PbSO_4 according to Hübener's method, slightly modified. Treated with 13 cc conc. HNO_3 , evaporated practically to dryness on the steam-bath, added 50 cc H_2O , and 0.5 cc Br and 2 cc of dilute HNO_3 (1:4). Heated, filtered, and washed with about 200 cc of hot water. Then precipitated with BaCl_2 .

Sulphur 0.73 0.80 0.57

These precipitates contained a little lead.

¹⁰ Kolb: *Dingl. pol. J.*, 209, p. 268. Ditte: *Ann. Chim. Phys.* [5], 14, p. 190.

From these determinations it seems quite certain that the larger part of the sulphur found as soluble sulphate under XII must have come from the solution and decomposition of lead sulphate first formed when the rubber was attacked by nitric acid.

At the suggestion of Dr. Hillebrand, four determinations were made of the amount of lead carried down with the barium sulphate precipitated in the usual way from the aqueous extract of the fusion mass. In spite of the presence of a large excess of sodium carbonate, some lead goes into solution. The preliminary treatment was according to XVIII, and 2 g of rubber instead of 0.50 g was taken each time.

XXXII. After fusion, the melts were dissolved in water. To each of the first two there was added 2 g of sodium bicarbonate in order to decompose any alkali plumbate. The solutions were heated on the steam-bath for one and one-half hours and then filtered from the insoluble. After acidifying with hydrochloric acid, barium sulphate was precipitated in the usual way.

	1	2	3	4
BaSO ₄ found (g)	0.5379	0.5395	0.5437	0.5417
Sulphur (per cent)	3.69	3.71	3.73	3.72

The barium sulphate precipitates were then mixed with soda and potash and fused. The melts were dissolved in water, filtered, and the residues washed with hot, very dilute sodium carbonate solution. The residues of barium carbonate and lead oxide were then dissolved in dilute nitric acid and the lead precipitated from the cold solutions by hydrogen sulphide. After standing over night in stoppered flasks, the precipitates of lead sulphide were filtered off, washed, dissolved in nitric acid, and finally converted into sulphate by evaporating down in porcelain crucibles with sulphuric acid and gently igniting.

	1	2	3	4
PbSO ₄ found (g)	0.0086	0.0071	0.0040	0.0045
Equivalent to BaSO ₄ (g)	0.0066	0.0055	0.0031	0.0035
Corrected BaSO ₄ (g)	0.5359	0.5379	0.5428	0.5407
Corrected sulphur (per cent)	3.68	3.69	3.73	3.71

It is quite evident from these figures that although notable quantities of lead sulphate are carried down with the barium sulphate, the correction in the percentage of sulphur is negligible.

The filtrates from the original precipitates were treated with hydrogen sulphide and gave slight precipitates. The alkaline filtrates from the barium carbonate and lead oxide were tested with ammonium sulphide and became brown. The next day there was a slight film of a dark color on the bottom of each of the beakers in which these solutions were tested. This was probably a mixture of small amounts of lead and iron sulphides. In all the solutions tested, as well as in the actual determinations of lead sulphate, greater amounts of lead were found in 1 and 2, which had been treated with bicarbonate. Apparently at the temperature of the steam-bath the lead bicarbonate probably formed was not decomposed.

CONCLUSIONS

Treatment of the rubber with nitric acid alone gives low results. (Compare XV with XVIII.) This is probably largely due to loss of free sulphur, since nitric acid alone does not completely oxidize sulphur to sulphuric acid in the length of time ordinarily taken for a determination.

The Hübener method can not be employed in the presence of mineral fillers which tend to form insoluble sulphates. This applies especially to barium carbonate and litharge.

A comparison of XX to XXVII shows that the fusion method gives results very close to those obtained by direct precipitation and by neutralization. The van't Kruys method gives high results.

The best results seem to be obtained by the use of method XVIII, according to which the rubber is decomposed by means of nitric acid saturated with bromine.

WASHINGTON, July 19, 1911.



LIBRARY OF CONGRESS



0 018 373 591 3